Evaluation of the Inhibitive Effect of Novel Quinoline Schiff Base on Corrosion of Mild Steel in HCl Solution

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Abstract—Schiff base (E)-2-methyl-N-(tetrazolo[1,5-a]quinolin-4-ylmethylene)aniline (QMA) was synthesized, and its inhibitive effect for mild steel in 1N HCl solution was investigated by weight loss measurement and electrochemical tests. From the weight loss measurements and electrochemical tests, it was observed that the inhibition efficiency increases with the increase in the Schiff base concentration and reaches a maximum at the optimum concentration. This is further confirmed by the decrease in corrosion rate. It is found that the system follows Langmuir adsorption isotherm.

Keywords—Schiff base, acid corrosion, electrochemical impedance spectroscopy, polarization.

I. INTRODUCTION

ACID solutions are used for the removal of undesirable scale &rust in many industrial processes. The use of inhibitors is one of the most practical methods to protect against excessive corrosion, especially in the acidic media [1]. The use of organic compounds containing oxygen, sulfur, and nitrogen to reduce corrosion attack on steel has been studied in detail [2]. Organic inhibitors generally protect the metal from corrosion by forming a film on the metal surface. The effectiveness is related to chemical composition, their molecular structure, and their affinities for the metal surface. The addition of corrosion inhibitors effectively protects the metal against an acid attack. Inhibitors are used in these processes to control metal dissolution [3]. Schiff base compounds are the condensation product of an amine and a ketone/aldehyde. Schiff base inhibitors have been reported as corrosion inhibitors for steel, copper and aluminum [4]-[8]. The adsorbed species protect the metal from corrosion. Adsorption depends on the nature, charge of the metal and the chemical structure of the inhibitor. The efficiency of Schiff’s bases is much higher than that of corresponding aldehyde and amines, and this may be due to the presence of a >C=N- group in the molecules. The planarity (π) and lone pairs of electrons present on N atoms are important structural features that determine the adsorption of Schiff’s base molecules on a metal surface. The aim of the present investigation was to synthesize and evaluate a novel Schiff base, (E)-2-methyl-N-(tetrazolo[1,5-a]quinolin-4-ylmethylene)aniline (QMA) and its performance as acid corrosion inhibitor for mild steel in 1M HCl by weight loss, polarization, and electrochemical impedance spectroscopy.

II. EXPERIMENTAL

A. Materials

Mild steel (MS) specimens having a composition of 0.09% P, 0.37% Si, 0.01% Al, 0.05% Mn, 0.19% C, 0.06% S, and the remainder Fe, were used for electrochemical polarization and impedance measurements. Specimens of 1cm × 1cm area were polished with emery paper number 120p1000, degreased with acetone, washed with double distilled water, and dried at room temperature before immersing in acid solution. The acid solutions were made from analytical grade 37% HCl and double-distilled water.

A Schiff base compound QMA with a structural formula as shown in Fig. 1 was synthesized to a stirred mixture of 2-chloro-quinoline-3-carbaldehyde (0.005 mol) and sodium azide (0.005 mol) in ethanol with a catalytic amount of PTSA and refluxed for 5p6 hours. After the reaction was complete, the reaction mixture was poured onto crushed ice; the solid mass thus separated out, was filtered, washed with distilled water, and dried. Then, tetrazolocarbaldehyde was treated with 2-methylaniline in ethanol with catalytic amount of sulphuric acid and refluxed for 5-6 hours. After the reaction was complete, the reaction mixture was poured onto crushed ice; the solid mass thus separated out, was filtered, washed with distilled water, and dried. Then, tetrazolocarbaldehyde was treated with 2-methylaniline in ethanol with catalytic amount of sulphuric acid and refluxed for 6-7 hours. After the reaction was complete, the reaction mixture was poured onto crushed ice; the solid mass thus separated out, was filtered, washed with distilled water, and dried.

Fig. 1 (E)-2-methyl-N-(tetrazolo[1,5-a]quinolin-4-ylmethylene) aniline (QMA)
B. Weight Loss Method

Weight loss measurements were carried out by weighing the prepared specimens before and after immersion for 4h in 100mL stagnant test solutions of 1N HCl in the presence and absence of various concentrations of QMA. From the weight loss data, percent inhibition efficiency (IE %) was calculated at different concentrations at 30°C. At the end of testing the specimens were carefully washed with acetone, dried, and then weighed.

C. Electrochemical Measurements

Polarization and impedance measurements were carried out in a conventional three electrode Cylindrical borosilicate glass cell consisting of a carbon steel working electrode(WE), a platinum counter electrode (CE), and saturated calomel electrode (SCE) as the reference electrode using a CH electrochemical analyzer model 608 C USA Potentiostat/ galvanostat. Polarization curves were obtained at a scan rate of 0.01V/s in the range of -250mV to +250mV vsEcorr. 1cm² surface area of a polished carbon steel specimen was exposed to 200mL of 1M HCl solution at 30°C, without and with various concentrations of inhibitor.

The Electrochemical impedance measurement was made in the frequency range of 0.1Hz to 1000Hz. The AC signal was 5mV peak to peak with 12 data points per decade. The double layer capacitance (Cdl) and the charge transfer resistance (Rct) were calculated from Nyquist plots.

III. RESULTS AND DISCUSSION

A. Weight Loss Measurements

Corrosion of MS in 1N HCl in the absence and presence of various concentrations (5-25 ppm) of QMA was determined by weight loss experiments. The corrosion rate (Wcorr) and the values of inhibition efficiency (IE %) were calculated according to following equation:

\[
IE\% = \frac{W_{corr} - W_{corr\ (inh)}}{W_{corr}} \times 100
\]

where \(W_{corr\ (inh)}\) and \(W_{corr}\) are the corrosion rates of MS with and without inhibitor respectively.

<table>
<thead>
<tr>
<th>Concentration of QMA in ppm</th>
<th>Corrosion Rate (mpy)</th>
<th>IE %</th>
</tr>
</thead>
<tbody>
<tr>
<td>1 N HCl</td>
<td>19.3</td>
<td>-</td>
</tr>
<tr>
<td>5</td>
<td>1.98</td>
<td>89.74</td>
</tr>
<tr>
<td>10</td>
<td>1.70</td>
<td>91.19</td>
</tr>
<tr>
<td>15</td>
<td>1.34</td>
<td>93.05</td>
</tr>
<tr>
<td>20</td>
<td>1.08</td>
<td>94.40</td>
</tr>
<tr>
<td>25</td>
<td>0.95</td>
<td>95.07</td>
</tr>
</tbody>
</table>

From Table I, it was found that the value of IE % increased with an increase in the concentration of QMA, confirming that the number of molecules adsorbed increased over the MS surface, blocking the active sites of acid attack and protecting the metal from corrosion. At 25 ppm of QMA added, the IE% was 95.07% and showed that QMA was a very effective inhibitor.

The nature of interaction between MS surface and inhibitor can be established by studying the type of isotherm for the system.

\[y = 1.033x + 0.527 \quad R^2 = 0.999\]

The inhibition efficiency increases with increase in the concentration of QMA. Thus QMA is adsorbed strongly on the MS surface forming a protective layer and shielding the metal from corrosion [9]-[11]. The observed inhibitive action of the QMA is attributed to the adsorption of the molecule on the metal surface by the interaction of the π electrons of the C=N group and the π electrons of the aromatic ring [12], [13]. Adsorption of QMA on the steel surface forms a barrier between metal and the aggressive environment. More over the presence of the electron donating groups on the Schiff base such as CH₃ increases the electron density on the nitrogen of C=N group resulted in the high inhibition efficiency. As concentration of QMA is increased, the fraction of steel covered by the adsorbed molecules, θ increases, leading to higher inhibition efficiency. Because of the direct relation between IE % and θ, the latter can be estimated using the relation \(θ = IE\% /100\). Fig. 2 shows the graphical relation between concentration C and ratio of concentration C and θ (C/θ) is a straight line with an approximate slope of unity. The system follows a Langmuir adsorption isotherm. The Langmuir isotherm is a linear relationship between C/θ and concentration C with a slope of ideally unity.

The adsorption behavior was described by using the Langmuir isotherm expressed as:

\[K C = θ / (1-θ)\]

where K is the equilibrium constant of the inhibitor adsorption process, C is the inhibitor concentration.

The free energy \(ΔG\) was calculated as follows:
\[ \Delta G_{\text{ads}} = -2.303RT \log(55.5K_{\text{ads}}) \]  

(3)

The negative value of \( \Delta G \) indicates a spontaneous adsorption of inhibitor molecules on the metal surface [14], [15].

**B. Polarization Measurement**

To evaluate the effect of QMA on the electrochemical behavior of MS, cathodic as well as anodic polarization studies were carried out as shown in Tafel plots in Fig. 3. The corrosion kinetic parameters such as corrosion potential \( E_{\text{corr}} \), corrosion current density \( I_{\text{corr}} \), and anodic and cathodic Tafel slopes \( b_a \) and \( b_c \) were derived from these curves and are given in Table II. The values of inhibition efficiency (IE\%) were calculated using the following equation:

\[ \text{IE} \% = \frac{I_{\text{corr}} - I'_{\text{corr}}}{I_{\text{corr}}} \times 100 \]  

(4)

where \( l_{\text{corr}} \) and \( l'_{\text{corr}} \) were the values of corrosion current densities without and with the different concentrations of QMA, respectively, which were determined by extrapolation of the cathodic and anodic Tafel lines to the corrosion potential \( E_{\text{corr}} \). It was observed from Table II that the values of IE\% increased from 94.18 to 98.29 with increase in the concentration of QMA, the highest IE\% was observed at 25 ppm concentration of QMA, and the \( I_{\text{corr}} \) values decreased from 0.2373 mA/cm\(^2\) to 0.0697 mA/cm\(^2\) with the increase in QMA concentration from 5 to 25 ppm, which confirmed the strong adsorption of QMA on the MS surface.

**TABLE II**

<table>
<thead>
<tr>
<th>Concentration of inhibitor (ppm)</th>
<th>(-E_{\text{corr}}) (V)</th>
<th>Tafel Constant (mV/decade)</th>
<th>( I_{\text{corr}} ) (mA/cm(^2))</th>
<th>( R_p ) (ohm cm(^2))</th>
<th>IE%</th>
<th>IE% Rp</th>
</tr>
</thead>
<tbody>
<tr>
<td>1 N HCl</td>
<td>0.508</td>
<td>b_a 145, b_c 135</td>
<td>4.0800</td>
<td>-</td>
<td>7</td>
<td>91.86</td>
</tr>
<tr>
<td>5</td>
<td>0.462</td>
<td>85, 105</td>
<td>0.2373</td>
<td>86</td>
<td>94.18</td>
<td>91.86</td>
</tr>
<tr>
<td>10</td>
<td>0.464</td>
<td>91, 106</td>
<td>0.2044</td>
<td>104</td>
<td>94.99</td>
<td>93.26</td>
</tr>
<tr>
<td>15</td>
<td>0.460</td>
<td>92, 108</td>
<td>0.1528</td>
<td>141</td>
<td>96.25</td>
<td>95.03</td>
</tr>
<tr>
<td>20</td>
<td>0.459</td>
<td>111, 136</td>
<td>0.0977</td>
<td>272</td>
<td>97.60</td>
<td>97.42</td>
</tr>
<tr>
<td>25</td>
<td>0.446</td>
<td>116, 151</td>
<td>0.0697</td>
<td>410</td>
<td>98.29</td>
<td>98.29</td>
</tr>
</tbody>
</table>

Fig. 3 Polarization curves for mild steel in 1N HCl in the absence and presence of different concentrations of QMA

The inhibiting property of QMA has also been evaluated by determining the polarization resistance \( R_p \) (ohm/cm\(^2\)). The corresponding polarization resistance \( R_p \) values for MS in 1 N HCl in the absence and presence of different concentrations of the QMA is given in Table II. The values of inhibition efficiency (IE\%) were calculated as follows:

\[ \text{IE} \% = \frac{R_{p(\text{inh})} - R_p}{R_p} \times 100 \]  

(5)

where \( R_p \) and \( R_{p(\text{inh})} \) were the polarization resistance in the absence and in the presence of the inhibitor respectively.

It was observed that the Tafel plot in 1 N HCl solutions for MS with different concentrations of QMA was similar (Fig. 2). Corrosion currents obtained with the additions of QMA was lower than the corrosion current obtained for MS in 1N HCl solution without the inhibitor. The IE % increased with an increase in the concentration of QMA (Table II). QMA has acted as an effective corrosion inhibitor suppressing both anodic and cathodic reactions due to their adsorption on the MS surface blocking the active sites [16]. Moreover, there was an anodic shift in \( E_{\text{corr}} \) values to show that the inhibitor may have retarded anodic reaction.

It was observed that the \( R_p \) values gradually increased with increase in the concentration of the inhibitor (Table II).
It can be seen from Table III, there was a gradual decrease in the value of $C_d$ with an increase in the concentration of QMA. The double layer between the charged metal surface and the solution is considered an electrical capacitor. The decrease of the capacitance with increase in the concentration of QMA can be related to the formation of a protective layer on the MS electrode surface. Inhibition efficiency (IE%) was found to increase with the increase in concentration of QMA. The results obtained from electrochemical impedance spectroscopy study have shown a similar trend as those obtained from electrochemical polarization and weight loss measurements.

The results suggest that the corrosion rate of MS was significantly decreased due to adsorption mechanism affecting both the anodic and cathodic processes.

IV. CONCLUSION

A novel Schiff base, QMA has shown good corrosion inhibiting characteristics for controlling corrosion of mild steel in 1N HCl. The inhibition efficiency increases with the increase in the inhibitor concentration and attains a maximum value at 25ppm. The inhibitor acted as a mixed type inhibitor controlling both cathodic and anodic reactions. The weight loss results, results of electrochemical impedance spectroscopy and polarization study are in good agreement. The adsorption of QMA on the metal surface follows Langmuir adsorption isotherm.

REFERENCES